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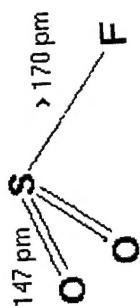
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## SYNOPSIS

The controversy between crystallographers and theoreticians concerning the structure of  $\text{SO}_2\text{F}^-$  has been resolved in favor of the theoreticians. Crystal structure determinations of the TAS, TAOS, TMA, and K salts of  $\text{SO}_2\text{F}^-$  show that the discrepancies between theory and experiment were due to severe oxygen/fluorine disorder in the crystals. Corrections for the disorder effects lead to a geometry that is in excellent agreement with the theoretical predictions and the observed vibrational spectra. The length and weakness of the S-F bond is confirmed by ab initio calculations and a normal coordinate analysis.

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Paul G. Watson, Tobias Borrman,  
Ashwani Vij, Jerry A. Boatz, Karl O. Christe\*

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The Structure of the  $\text{SO}_2\text{F}^-$  Anion,  
a Problem Case

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# The Structure of the $\text{SO}_2\text{F}^-$ Anion, a Problem Case<sup>1</sup>

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## Abstract

Recently, room-temperature crystal structures of  $\text{SO}_2\text{F}^-$  in its  $\text{K}^+$  and  $\text{Rb}^+$  salts were published in *Z. Anorg. Allg. Chem.* **1999**, *625*, 385 and claimed to represent the first reliable geometries for  $\text{SO}_2\text{F}^-$ . However, their almost identical S-O and S-F bond lengths and O-S-O and O-S-F bond angles are in sharp contrast to the results from theoretical calculations. To clarify this discrepancy, the new  $[(\text{CH}_3)_2\text{N}]_3\text{SO}^+$  and the known  $[\text{N}(\text{CH}_3)_4]^+$ ,  $[(\text{CH}_3)_2\text{N}]_3\text{S}^+$  and  $\text{K}^+$  salts of  $\text{SO}_2\text{F}^-$  were prepared and their crystal structures studied at low temperatures. Furthermore, the results from previous RHF and MP2 calculations were confirmed at the RHF, B3LYP and CCSD(T) levels of theory using different basis sets. It is shown that all the  $\text{SO}_2\text{F}^-$  salts studied so far exhibit varying degrees of oxygen-fluorine and, in some cases, oxygen site disorders, with  $[(\text{CH}_3)_2\text{N}]_3\text{SO}^+$   $\text{SO}_2\text{F}^-$  at 113 °K showing the least disorder with  $r(\text{SF})-r(\text{SO}) = 17 \text{ pm}$  and  $\langle (\text{OSO})- \rangle - \langle (\text{FSO})- \rangle = 6^\circ$ . Refinement of the disorder occupancy factors and extrapolation of the observed bond distances for zero disorder resulted in a geometry very close to that predicted by theory. The correctness of the theoretical predictions for  $\text{SO}_2\text{F}^-$  is further supported by the good agreement between the calculated and the experimentally observed

vibrational frequencies and their comparison with those of isoelectronic  $\text{ClO}_2\text{F}$ . A normal coordinate analysis of  $\text{SO}_2\text{F}^-$  confirms the weakness of the S-F bond with a stretching force constant of only 1.63 mdyn/Å and shows that there is no highly characteristic S-F stretching mode. The S-F stretch strongly couples with the  $\text{SO}_2$  deformation modes and is concentrated in the two lowest  $a'$  frequencies.

## Introduction

The  $\text{SO}_2\text{F}^-$  anion has been known since 1953 from the work of Seel and coworkers.<sup>2-5</sup> Its vibrational spectra were studied by Seel and Boudier,<sup>6</sup> Paetzold and Aurich,<sup>7</sup> Robinson and coworkers,<sup>8</sup> Moock and coworkers,<sup>9</sup> and Kornath and coworkers<sup>10</sup> and correctly assigned with the help of ab initio calculations in terms of a  $C_s$  symmetry structure with a predicted geometry of  $r(\text{SF}) = 170$  pm,  $r(\text{SO}) = 146$  pm,  $\angle(\text{OSO}) = 113.2^\circ$ , and  $\angle(\text{OSF}) = 100.6^\circ$ .<sup>10</sup> The agreement between observed and calculated vibrational frequencies was good and supported the calculated geometry. Furthermore, the calculated geometry of  $\text{SO}_2\text{F}^-$  is similar to that experimentally observed for isoelectronic  $\text{ClO}_2\text{F}$ .<sup>11</sup>

The geometry, predicted from the theoretical calculation and supported by the vibrational analysis<sup>10</sup> is in conflict with the results from three recent X-ray diffraction studies. Zhu and coworkers reported the crystal structure of  $\text{Ph}_3\text{PCF}_2\text{H}^+\text{SO}_2\text{F}^-$ .<sup>12</sup> In this structure, the  $\text{SO}_2\text{F}^-$  anion is both fluorine-oxygen and oxygen-site disordered and its geometry was listed as  $r(\text{S-F}) = 151.6(6)$  pm,  $r(\text{S-O1}) = 143.6$  pm,  $r(\text{S-O2}) = 141(1)$  pm,

$\angle(F-S-O1) = 110.5(5)^\circ$ . The second structure was obtained by Kuhn and coworkers for 2-fluoro-1,3-diisopropyl-4,5-dimethylimidazolium fluorosulfite.<sup>13,14</sup> In this structure, the  $SO_2F^-$  anion is similarly disordered and, therefore, the authors considered a discussion of the  $SO_2F^-$  geometry inappropriate.<sup>13</sup> In the most recent study by Kessler and Jansen,<sup>15</sup> the room temperature structures of  $KSO_2F$  and  $RbSO_2F$  were given with  $r(SF) = 159.1(2)$  pm,  $r(SO) = 152.6(2)$  pm,  $\angle(OSO) = 104.9(2)^\circ$ , and  $\angle(OSF) = 102.8(1)^\circ$  for the  $K^+$  salt. Possible disorder was ignored and the obtained structural parameters were taken as the correct geometry of isolated  $SO_2F^-$ , concluding that the previously published,<sup>10,14</sup> theoretically calculated structures were invalid. This conclusion, however, raises serious questions. Above all, how could a geometry which deviates by 18 pm for  $r(SF)$  and by 8° for  $\angle(OSO)$  from the true structure duplicate well the observed vibrational frequencies?

In this paper we wish to report the preparation of the  $(Me_2N)_3S^+$  (TAS),<sup>16</sup>  $(Me_2N)_3SO^+$  (TAOS),  $N(CH_3)_4^+$  (TMA),<sup>3,4</sup> and  $K^+$  salts of  $SO_2F^-$  and their low-temperature crystal-structures. It is shown that the true geometry of  $SO_2F^-$  is close to the calculated ones and that the varying degrees of deviation, exhibited in the crystal structures of the different salts, are due to the propensity of  $SO_2F^-$  to undergo oxygen-fluorine and, in some cases, oxygen-site disorders.

## Experimental Section

Standard inert atmosphere techniques were used for the manipulation of all reagents and reaction products. Infrared spectra were recorded on a Nicolet DX-55-FT-IR spectrometer using Nujol/Kel-F mulls, NMR-spectra on a Bruker WP80SY and referenced to  $\text{Me}_4\text{Si}$  ( $^1\text{H}$ ) and  $\text{CFCl}_3$  ( $^{19}\text{F}$ ). The NMR samples, dissolved in liquid  $\text{SO}_2$ , were contained in sealed 5 mm glass tubes. Commercial grade solvents ( $\text{MeCN}$ ,  $\text{Et}_2\text{O}$ ) were dried and purified by employing conventional procedures.<sup>17</sup> Commercial  $\text{SO}_2$  was stored over  $\text{P}_4\text{O}_{10}$  at room temperature. TAOSF was obtained, similarly to TASF, from  $\text{OSF}_4$  and  $\text{Me}_3\text{SiNMe}_2$ ;<sup>18</sup> TASF<sup>16,19</sup> and TMAF<sup>20</sup> were prepared as described, and the fluorosulfites were synthesized according to the method reported for  $\text{TAS}^+\text{SO}_2\text{F}^-$ .<sup>16</sup> The spectroscopic data of  $\text{TAS}^+\text{SO}_2\text{F}^-$  (1) and  $\text{TMA}^+\text{SO}_2\text{F}^-$  (3) agreed with those reported in the literature.<sup>3,4,16</sup>

*at -30 °C  
try to keep  
mL*  
 $\text{TAOS}^+\text{SO}_2\text{F}^-$  (2) was obtained by dissolving  $\text{TAOS}^+\text{Me}_3\text{SiF}_2^-$  (1.7 mmol) in 10  
volatile material was removed under vacuum. 2 (0.45 g, 100% yield) remained as a  
colorless solid, mp 239 °C (dec). *remove space*

Salts 1–3 were recrystallised from  $\text{MeCN}/\text{Et}_2\text{O}$  mixtures at –40 °C to give single crystals suitable for X-ray structure determinations.

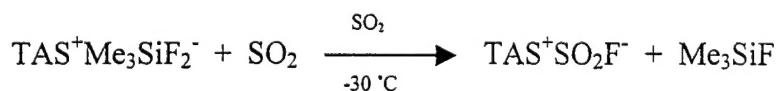
**Crystal Structure Determinations.** The crystals were mounted on thin glass fibers using the oil-drop technique (Kel-F oil). The data were collected on a Siemens P4 diffractometer using Mo-K $\alpha$  radiation ( $\lambda = 71.073$  pm) at the given temperatures. The

structures were solved by direct methods. Hydrogen atoms were located from difference electron density maps and refined isotropically.

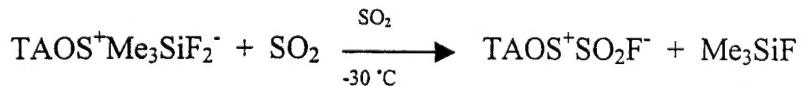
**Computational Methods.** The geometries and harmonic vibrational frequencies of  $\text{SOF}_2$ ,  $\text{SO}_2\text{F}_2$ , and  $\text{SO}_2\text{F}^-$  were calculated using restricted Hartree-Fock (RHF), density functional theory (using the B3LYP hybrid functional<sup>21</sup>), and coupled-cluster methods (CCSD(T)<sup>22</sup>). The 6-311+G(2d)<sup>23</sup> and the augmented correlation-consistent polarized valence double- and triple-zeta basis sets<sup>24</sup> were used. All calculations were performed using GAMESS<sup>25</sup> and GAUSSIAN 94.<sup>26</sup>

## Results and Discussion

**Synthesis and Properties of  $\text{TAOS}^+\text{SO}_2\text{F}$ .** TAS fluoride,  $(\text{Me}_2\text{N})_3\text{S}^+\text{Me}_3\text{SiF}_2^-$ ,<sup>19</sup> is an excellent fluoride ion donor which readily transfers a fluoride ion to the more acidic  $\text{SO}_2$  molecule.<sup>16</sup>



Replacement of TASF by TAOSF,  $(\text{Me}_2\text{N})_3\text{SO}^+\text{Me}_3\text{SiF}_2^-$ , results in the corresponding  $\text{TAOS}^+\text{SO}_2\text{F}^-$  salt in quantitative yield.



The new  $\text{TAOS}^+\text{SO}_2\text{F}^-$  salt is a colorless solid that is stable at room temperature and melts at 239 °C with decomposition. It was characterized by its crystal structure (see below).

**Crystal Structures of  $\text{TAS}^+\text{SO}_2\text{F}^-$  (1),  $\text{TAOS}^+\text{SO}_2\text{F}^-$  (2),  $\text{TMA}^+\text{SO}_2\text{F}^-$  (3), and  $\text{K}^+\text{SO}_2\text{F}^-$  (4).** The crystal structures of **1–4** were determined at 173 °K and those of **2** and **4** also at 113 °K. The crystal and structure refinement data and the bond distances and angles of the  $\text{SO}_2\text{F}^-$  anions are given in Tables 1 and 2, respectively. The molecular units with atom labeling, the packing diagrams and the closest anion-cations interactions are shown in Figures 1–7. As can be seen, compounds **1–4** are predominantly ionic, containing isolated  $\text{SO}_2\text{F}^-$  anions.

The refinement of the  $\text{TAS}^+\text{SO}_2\text{F}^-$  (1) structure was not trivial. Initial structure refinements in the space group Pnma put the atoms S, F and O(2) on special positions located on a mirror plane, while O(1) was site-disordered occupying split positions off the plane with occupancy factors of one half. In addition, the similarity of the S-O(2) and S-F bond lengths was also indicative of fluorine/oxygen disorder. This solution resulted in an apparent unusually short S-O(1) bond of ~140 pm, about 5 pm shorter than our theoretical predictions for  $r(\text{S-O})$ , with large thermal motions around the  $\text{SO}_2\text{F}$  atoms. Since the theoretical calculations also predict that the S-F bonds in  $\text{SO}_2\text{F}^-$  are much longer than the S-O bonds, oxygen/fluorine disorder could only lengthen but not shorten the S-O bonds. The same argument applies to libration corrections. Therefore, this refinement model does not result in a plausible structure for  $\text{SO}_2\text{F}^-$ . The shape of the thermal ellipsoids of the atoms in the  $\text{SO}_2\text{F}^-$  anion suggested that these atoms might be disordered across the crystallographic mirror plane. Consequently, the positions of all

atoms of  $\text{SO}_2\text{F}^-$  were fixed off this plane, resulting in two sets of disordered anions with half occupancy (see Figure 1). The resulting refinement gave a plausible value of 146.5(3) pm for the S-O(2) bond distance, indicating that O(2) is not affected by oxygen/fluorine disorder, and resulted in slightly improved R factors. However, the O(1) and F positions are disordered and, therefore, the S-O and S-F bond distances and their bond angles could not be uniquely determined in this manner. The disorder was first modeled with equal O(1)/F occupancy factors. However, it was found that the occupancy factors cannot be equal because the two resulting S-O(1)/F bond lengths differed by 3.7 pm. The shortcomings of this model are also apparent from the sum of the OSO and OSF bond angles which total 331.6 ° instead of the theoretically predicted 314 °. To obtain the correct occupancy factors which can account for the different S-O(1)/F bond lengths, the sums of the individual occupancy factors were restrained to their required totals (1.0 and 1.0), and the variable individual occupancy factors were refined to give the best fit with the observed data. The occupancy factors of 2-4 were refined in the same manner, and the results are summarized in Table 2.

The above analysis indicates that the disorder of the  $\text{SO}_2\text{F}^-$  anion involves a symmetry plane defined by the S and O(2) atoms and the free valence electron pair on S, with O(1) and F being located off this plane. Therefore, only O(1) and F are affected by this disorder and the actual bond length of the S-O bond is well determined by the observed S-O(2) distance.

The packing diagram of  $\text{TAS}^+ \text{SO}_2\text{F}^-$  (Fig. 2) shows a three-dimensional network of intermolecular H...O/F contacts. Out of the four potential binding sites in the disordered  $\text{SO}_2\text{F}^-$  anion, the F/O atoms form a pseudo ten-membered ring (Fig. 1) with a

H1A...F/O distance of 250(3) pm. We use the term "pseudo ring", because one side of the ring structure is always left open due to the disorder-induced half-occupancy. These pseudo ring structures are then connected to the others in the same plane *via* the O2/O2A contacts i.e., H1C...O2 and H2C...O2 at 258(3) and 266(2) pm, respectively. This network is then linked to the other networks via bonding from O1/F1 at 253(3) pm.

The disorder of the  $\text{SO}_2\text{F}^-$  anion in the **TAOS salt** is less complicated than that in the TAS salt because it exhibits only the O(1)/F disorder with respect to the mirror plane formed by O(2), S, and the free valence electron pair on sulfur, but not the additional disorder with respect to the crystallographic mirror plane. Furthermore, from all the presently known  $\text{SO}_2\text{F}^-$  structures, that of the TAOS salt at 113 °K (Figures 3 and 4) exhibits the least amount of oxygen/fluorine disorder, resulting in the largest differences between the apparent S-F and S-O bond lengths and O-S-O and F-S-O bond angles. The degree of O/F disorder also decreases with decreasing temperature, as was established by carrying out the structure determinations at 173 and 113 °K. As in the case of the TAS salt, the S-O(2) bond, which lies on the molecular symmetry plane, is not affected by disorder, resulting in an S-O(2) bond length of 146.8(3) pm which is, within experimental error, identical to that found for the unique oxygen in the TAS salt and, therefore, should be representative of the true S-O bond length in ordered  $\text{SO}_2\text{F}^-$ . Furthermore, the sum of the OSO and OSF bond angles amounts to 313.9 ° which is in excellent agreement with the value of 314 °, predicted from the ab initio calculations. The facts that  $r(\text{S-O}\{1\})$  and  $r(\text{S-F})$  in the 113 °K structure of  $\text{TAOS}^+\text{SO}_2\text{F}^-$ , obtained by refinement assuming an ordered structure, are still somewhat longer and shorter, respectively, than the theoretical

predictions for free  $\text{SO}_2\text{F}^-$  and that  $r(\text{S}-\text{O}(1))$  is 1.6 pm longer than  $r(\text{S}-\text{O}(2))$ , indicates that even in this case there is still some remaining disorder of O(1) and fluorine.

Due to the absence of strong disorder in  $\text{TAOS}^+\text{SO}_2\text{F}^-$  at 113 °K, its crystal packing (Fig. 4) can be discussed more meaningfully. The  $\text{CH}_3$  group of C3 bridges to two  $\text{SO}_2\text{F}^-$  anions via the  $\text{F}1\ldots\text{H}3\text{A}-\text{C}3-\text{H}3\text{C}\ldots\text{O}1$  links, forming a zig-zag chain ( $\text{F}1\ldots\text{H}3\text{A} = 243(3)$  pm,  $\text{O}1\ldots\text{H}3\text{C} = 255(4)$  pm). These chains then form a 3-dimensional network by the formation of an intra-molecular bifurcated fluorine contact ( $\text{H}1\text{A}\ldots\text{F}1$ , 246(4) pm). These three contacts are the shortest long-range contact distances<sup>27</sup> in the crystal lattice.

Refinement of the 173 °K structure of  $\text{TAOS}^+\text{SO}_2\text{F}^-$ , assuming no disorder, suffered from shortcomings similar to those previously reported in the literature for the  $\text{Ph}_3\text{PCF}_2\text{H}^+$  salt<sup>12</sup> and provided no useful information.

In the  $\text{TMA}^+\text{SO}_2\text{F}^-$  salt (see Figures 5 and 6 and Tables 1 and 2) at 173 °K, the  $\text{SO}_2\text{F}^-$  anion shows, as in the  $\text{TAOS}^+$  salt, only O(1)/F disorder with respect to the mirror plane formed by O(2), S, and the free valence electron pair on sulfur, resulting again in a unique S-O(2) distance of 147.0(2) pm, which is almost identical to those found for the  $\text{TAS}^+$  and  $\text{TAOS}^+$  salts. Refinements, carried out assuming equal occupancy factors for O(1) and F, resulted again in different bond lengths for the two S-O(1)/F bonds, demonstrating the need for refining the occupancy factors.

The packing diagram for  $\text{TMA}^+\text{SO}_2\text{F}^-$  is shown in Figure 6 and exhibits bifurcated bridges between O(2) and two hydrogen atoms from the same TMA cation, and between one of the O(1)/F atoms and hydrogen atoms from two different TMA cations, with the other O(1)/F atom forming a single bridge to a TMA cation. The bifurcated bridging is

similar to that observed for  $\text{TAOS}^+\text{SO}_2\text{F}^-$ , resulting in 16-membered rings that contain two cations and two anions and are three-dimensionally cross-linked to other rings. The two shortest of the many close distances are  $\text{H2C...O1/F1}$  at  $235(2)$  pm and  $\text{H1B...O2}$  at  $243(2)$  pm.

The **low-temperature (113 °K) structure of  $\text{K}^+\text{SO}_2\text{F}^-$**  is shown in Figure 7. This structure also exhibits oxygen/fluorine disorder but, contrary to the TAS, TAOS, and TMA salt structures, both oxygen atoms participate in the disorder. Therefore, this structure does not provide a unique S-O bond distance. The differences between the  $\text{KSO}_2\text{F}$  and the TAS, TAOS and TMA salt structures suggest a different kind of disorder. In the TAS, TAOS, and TMA salts, which exhibit significant anion-cation interactions through hydrogen bridges, the disorder is related to a symmetry plane through O(2), S, and the free valence electron pair on sulfur, while in the  $\text{K}^+$  salt, which exhibits weaker anion-cation bridging, it appears to involve positional disorder around a pseudo three-fold axis along sulfur and its free valence electron pair. In the positionally disordered structure of  $\text{KSO}_2\text{F}$ , the occupancy factors for the three O/F positions are not equal and one bond is considerably longer than the other two, requiring again a refinement of the occupancy factors. Except for the expected temperature effects, our low-temperature structure of  $\text{K}^+\text{SO}_2\text{F}^-$  is essentially identical to that previously found at room temperature,<sup>15</sup> although the longer S-F/O bond is about 4 pm longer at the lower temperature, indicating a decrease in disorder with decreasing temperature.

As can be seen from Figure 7, the potassium ion is surrounded by 6  $\text{SO}_2\text{F}^-$  anions. Out of these, 3 form bidentate and 3 form monodentate bridges to  $\text{K}^+$ . Out of the 3 bidentate  $\text{SO}_2\text{F}^-$  groups, one is bridged through the two atoms with the higher oxygen

occupancy factors, designated as O/F; the other two are bridged through one atom with the higher oxygen occupancy factor (O/F) and one atom with the higher fluorine occupancy factor, designated as F/O. Out of the 3 monodentate  $\text{SO}_2\text{F}^-$  groups, 2 are bridged through O/F atoms and one through F/O atoms, resulting in an unusual coordination number of 9 around  $\text{K}^+$ .

The above structure analyses resulted for the TAS, TAOS, and TMA salts in a well defined S-O(2) distance of about 147 pm that is in excellent agreement with the theoretical predictions (see below). However in all these structures, the observed S-F distances are too short due to O(1)/F disorder with varying occupancy factors. Since the observed bond lengths should be a function of the corresponding occupancy factors, one should be able to extrapolate to a 100 % occupancy factor and thus obtain an estimate for the individual S-F bond length in  $\text{SO}_2\text{F}^-$ . This was done in Figure 8, which shows that the length of the S-F bond should exceed 170 pm, in agreement with the theoretical predictions. The previous conclusion<sup>15</sup> that the geometry derived from the strongly disordered room temperature structure of  $\text{KSO}_2\text{F}$  represents the first reliable geometry for  $\text{SO}_2\text{F}^-$  is obviously flawed. The pronounced tendency of  $\text{SO}_2\text{F}^-$  to undergo disorder can be attributed to its pyramidal shape, the similar size and electronegativity of its ligands, and packing arrangements that favor disorder.

## Computational Results

Molecular vibrations occur on a very fast time scale and, therefore, are not affected by disorder phenomena in the crystal. Furthermore, the vibrational frequencies of a molecule are uniquely determined by its geometry. Therefore, good agreement

between observed and calculated frequencies can be obtained only with the correct geometry,<sup>22</sup> and not, as previously implied,<sup>15</sup> with an incorrect geometry. Good quality theoretical calculations generally yield relatively accurate bond angles.<sup>22</sup> Depending on the level of theory and the quality of the basis set used, the bond distances are generally also accurate within a few pm. When comparing calculated and observed frequencies, it should be kept in mind that the calculated values are harmonic frequencies while the observed ones are anharmonic values, thus giving rise to small discrepancies. Furthermore, different physical states can cause minor deviations. The calculated values are for the isolated free species in the gas phase, while observed spectra are frequently for solids or liquids. However, these effects are generally small, and systematic over- or under- estimation of bond lengths and vibrational frequencies at certain levels of theory can be corrected by using scaling factors.

Reliable experimental vibrational spectra for SO<sub>2</sub>F<sup>-</sup> have recently been published by Kornath and coworkers and were shown to be in reasonable agreement with the values calculated at the uncorrelated RHF/6-31 +G\* level of theory.<sup>10</sup> The calculated geometry was also in good agreement with values<sup>14</sup> previously obtained at the RHF/6-31+G(3df) and MP2/6-31+G(3df) levels of theory. We have calculated the geometry and vibrational frequencies of SO<sub>2</sub>F<sup>-</sup> at the uncorrelated RHF, the correlated CCSD(T), and the density functional B3LYP levels of theory using 6-311+G(2d),<sup>21</sup> aug-cc-pvdz, and aug-cc-pvtz<sup>22</sup> basis sets. To judge the reliability of these calculations, the geometries and vibrational frequencies of the closely related and well known SOF<sub>2</sub><sup>28,29</sup> and SO<sub>2</sub>F<sub>2</sub><sup>30,31</sup> molecules were also calculated.

The results from these calculations are summarized in Tables 3 and 4, respectively, and show that the bond angles of free  $\text{SO}_2\text{F}^-$  change only little with the method of calculation and the basis set. Values of  $\angle(\text{FSO}) = 100.5^\circ$  and  $\angle(\text{OSO}) = 113^\circ$  should be close to the actual values. The calculated S-O bond length range in  $\text{SO}_2\text{F}^-$  is also narrow, and this bond is predicted to be about 147 pm which is in excellent agreement with the value of 146.8(3) pm observed in one of our crystal structure studies. The S-F bond length, however, is more sensitive to the level of theory and quality of the basis set chosen and is predicted to fall into the range of 170-176 pm, which is not too different from the minimum bond length of 165.7 pm, derived from our crystal structure studies. This relatively wide range is not surprising since the S-F bond in  $\text{SO}_2\text{F}^-$  is highly polar (see normal coordinate analysis), and the calculations for highly polar bonds are very sensitive to correlation and basis set polarization functions. Based on our predictions (see Table 3), the difference between  $r(\text{SF})$  and  $r(\text{SO})$  should range from 23 to 29 pm, which is in marked contrast to the 6 pm, required if the published<sup>15</sup> crystal structures of  $\text{KSO}_2\text{F}$  and  $\text{RbSO}_2\text{F}$  were representative of the free  $\text{SO}_2\text{F}^-$  anion. Tables 3 and 4 also show that even the RHF calculations with a good basis set give surprisingly good results for these sulfur oxofluorides. For  $\text{OSF}_2$ , which has the same  $C_s$  symmetry as  $\text{SO}_2\text{F}^-$ , only the best coupled cluster calculation, CCSD(T)/aug-cc-pvtz, gave a better result, i.e., a smaller average frequency deviation and smaller scaling factors, than the RHF/6-311+G(2d) calculation.

Additional evidence for the existence of a long and highly polar S-F bond in  $\text{SO}_2\text{F}^-$  comes from the structure<sup>11</sup> and vibrational spectra<sup>32</sup> of isoelectronic  $\text{ClO}_2\text{F}$ . The observed difference between the bond lengths of Cl-F and Cl=O is 27.4 pm,<sup>11</sup> and the

structure is very similar to that predicted for free gaseous  $\text{SO}_2\text{F}^-$  (see Table 3). Further support for the geometry, predicted by us in Table 3 for free  $\text{SO}_2\text{F}^-$ , comes from a recent study of  $\text{N}(\text{CH}_3)_4^+\text{SO}_2\text{CN}^-$ .<sup>33</sup> The observed vibrational spectra were found to be in good agreement with those calculated at the RHF/6-31+G\* level for a minimum energy structure with  $r(\text{SO}) = 146.7$  pm,  $r(\text{SC}) = 190.8$  pm,  $\angle(\text{OSO}) = 114.0^\circ$ , and  $\angle(\text{OSC}) = 100.8^\circ$ . These structural parameters are very similar to those predicted by us for  $\text{SO}_2\text{F}^-$  in Table 3.

### Normal Coordinate Analysis

The assignment and force constant of the S-F stretching mode of  $\text{SO}_2\text{F}^-$  are of significant interest. If indeed the geometry proposed in reference 15 for  $\text{SO}_2\text{F}^-$  in its  $\text{K}^+$  and  $\text{Rb}^+$  salts with  $r(\text{SF}) = 159$  pm were correct, its S-F stretching mode and force constant should be comparable to those found for  $\text{SF}_2$  [ $r_e(\text{SF}) = 159.2$  pm,  $\nu_{\text{as}} \text{SF}_2 = 813.0 \text{ cm}^{-1}$ ,  $\nu_{\text{sym}} \text{SF}_2 = 838.5 \text{ cm}^{-1}$ ,  $f_r = 4.72 \text{ mdyn}/\text{\AA}$ ].<sup>34,35</sup> If, on the other hand,  $r(\text{SF})$  in  $\text{SO}_2\text{F}^-$  falls into the range of 170-176 pm, its stretching mode and force constant should be much lower.

The assignment of the S-F stretching mode in  $\text{SO}_2\text{F}^-$  has been controversial from the very beginning. Paetzold and Aurich published the infrared and Raman spectra of  $\text{KSO}_2\text{F}$  in 1965 and assigned the  $595 \text{ cm}^{-1}$  band to the  $\text{SO}_2$  scissoring mode and the  $496 \text{ cm}^{-1}$  band to the S-F stretching mode.<sup>7</sup> They attributed the surprisingly low S-F stretching frequency to a highly polar S-F bond and to fluorine bridging between the  $\text{SO}_2\text{F}^-$  ions. Shortly afterwards, Seel and Boudier reported the infrared spectra of  $\text{KSO}_2\text{F}$ ,  $\text{RbSO}_2\text{F}$  and  $\text{CsSO}_2\text{F}$ . Since the frequencies of the bands at about  $595$  and  $495 \text{ cm}^{-1}$  changed only little for the different cations, they correctly concluded that the  $\text{SO}_2\text{F}^-$  anions should not

be significantly associated. In view of the lack of significant fluorine bridging which could have explained the low S-F stretching frequency, they proposed to assign the 595  $\text{cm}^{-1}$  band to the S-F stretching mode.<sup>6</sup> In a 1990 publication on  $\text{NH}_4^+\text{SO}_2\text{F}^-$ , Moock and coworkers followed<sup>9</sup> the assignments of Paetzold and Aurich,<sup>7</sup> whereas in 1997 Kornath and coworkers followed<sup>10</sup> that of Seel and Boudier.<sup>6</sup> In the most recent paper,<sup>15</sup> Kessler and Jansen believed to have confirmed the assignments of Seel and Boudier, because they observed isolated  $\text{SO}_2\text{F}^-$  ions in the crystal structures of  $\text{KSO}_2\text{F}$  and  $\text{RbSO}_2\text{F}$  and falsely assumed that Kornath's ab initio calculations had established not only the observed frequencies, but also their identities. The identification of fundamental vibrations within a given symmetry species, however, requires the knowledge of the potential energy distribution (PED) obtainable through a standard normal coordinate analysis.

The results from our normal coordinate analysis of  $\text{SO}_2\text{F}^-$  are summarized in Table 5. The unscaled CCSD(T)/aug-cc-pvtz data were used because they duplicate the observed frequencies well and require by far the smallest scaling factors. The PED of Table 5 contains a big surprise. It shows that  $v_2$ ,  $v_3$ , and  $v_4$  are all strongly mixed and that  $S_2$ , the symmetry coordinate of the S-F stretch, contributes only 20 % to  $v_2$ , and 40 and 39 % to  $v_3$  and  $v_4$ , respectively. Thus, the S-F stretch is distributed over all three fundamental vibrations and is concentrated in  $v_3$  and  $v_4$ . An inspection of the signs in the PED reveals the following coupling effects. 580  $\text{cm}^{-1}$ : in phase coupling of  $\delta_{\text{sym}} \text{FSO}_2$  (45%) with  $\delta_{\text{sciss}} \text{SO}_2$  (31%), and their out of phase coupling with vSF (20%); 506  $\text{cm}^{-1}$ : in phase coupling of  $\delta_{\text{sciss}} \text{SO}_2$  (56%) with vSF (40%), and their out of phase coupling with  $\delta_{\text{sym}} \text{FSO}_2$  (3%); 365  $\text{cm}^{-1}$ : in phase coupling of  $\delta_{\text{sym}} \text{FSO}_2$  (51%) with vSF (39%), and

their out of phase coupling with  $\delta_{\text{sciss}}$   $\text{SO}_2$  (10%). This analysis demonstrates the irrelevance and fallacies of arguments over the assignment of certain modes without the benefits of a normal coordinate analysis. As so often, nature disregards our desire to paint simple black and white pictures of complex issues.

A second surprising result from the normal coordinate analysis is the very small value of 1.6 mdyn/Å for the SF stretching force constant. Normal, predominantly covalent S-F bonds exhibit values ranging from about 4.5 to 5.4 mdyn/Å.<sup>36,37</sup> The low value of the SF stretching force constant demonstrates the high polarity of the S-F bond in  $\text{SO}_2\text{F}^-$ . Furthermore, the 39% contribution of S-F stretching to the low frequency 365  $\text{cm}^{-1}$  mode provides a dissociative mode with a low activation energy barrier toward the loss of a fluoride ion.

### Conclusions

- (i) The experimental crystal structures of the  $\text{TMA}^+$ ,  $\text{TAS}^+$ ,  $\text{TAOS}^+$  and  $\text{K}^+$  salts of  $\text{SO}_2\text{F}^-$  show that all the crystal structures reported so far for  $\text{SO}_2\text{F}^-$  suffer from severe oxygen/fluorine disorder and that the given geometries do not reflect the true structure of  $\text{SO}_2\text{F}^-$ . In the structures of the  $\text{TMA}^+$ ,  $\text{TAS}^+$ , and  $\text{TAOS}^+$  salts, one oxygen atom is ordered and its S-O distance of 147 pm is well defined. The S-F distances can be estimated to equal or exceed 170 pm by refining the occupancy factors in the O/F disorders and by an extrapolation of a correlation between the occupancy factors and the observed bond lengths to 100 % occupancy.

**Supporting Information Available:** Tables of structure determination summaries, atomic coordinates, bond lengths and angles, anisotropic displacement parameters, and hydrogen coordinates of (1), (2), (3) and (4) in CIF format. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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## Diagram Captions

Figure 1: Crystal structure, numbering scheme, and hydrogen bridging of  $\text{TAS}^+\text{SO}_2\text{F}^-$  (1), showing the two disordered positions of  $\text{SO}_2\text{F}^-$ . The displacement ellipsoids are drawn at the 50 % probability level, and the hydrogen atoms were located from difference electron density maps.

Figure 2: Packing diagram and hydrogen bridging of  $\text{TAS}^+\text{SO}_2\text{F}^-$  (1).

Figure 3: Crystal structure, numbering scheme, and hydrogen bridging of  $\text{TAOS}^+\text{SO}_2\text{F}^-$  (2) at 113 °K. The displacement ellipsoids are drawn at the 50 % probability level, and the hydrogen atoms were located from difference electron density maps.

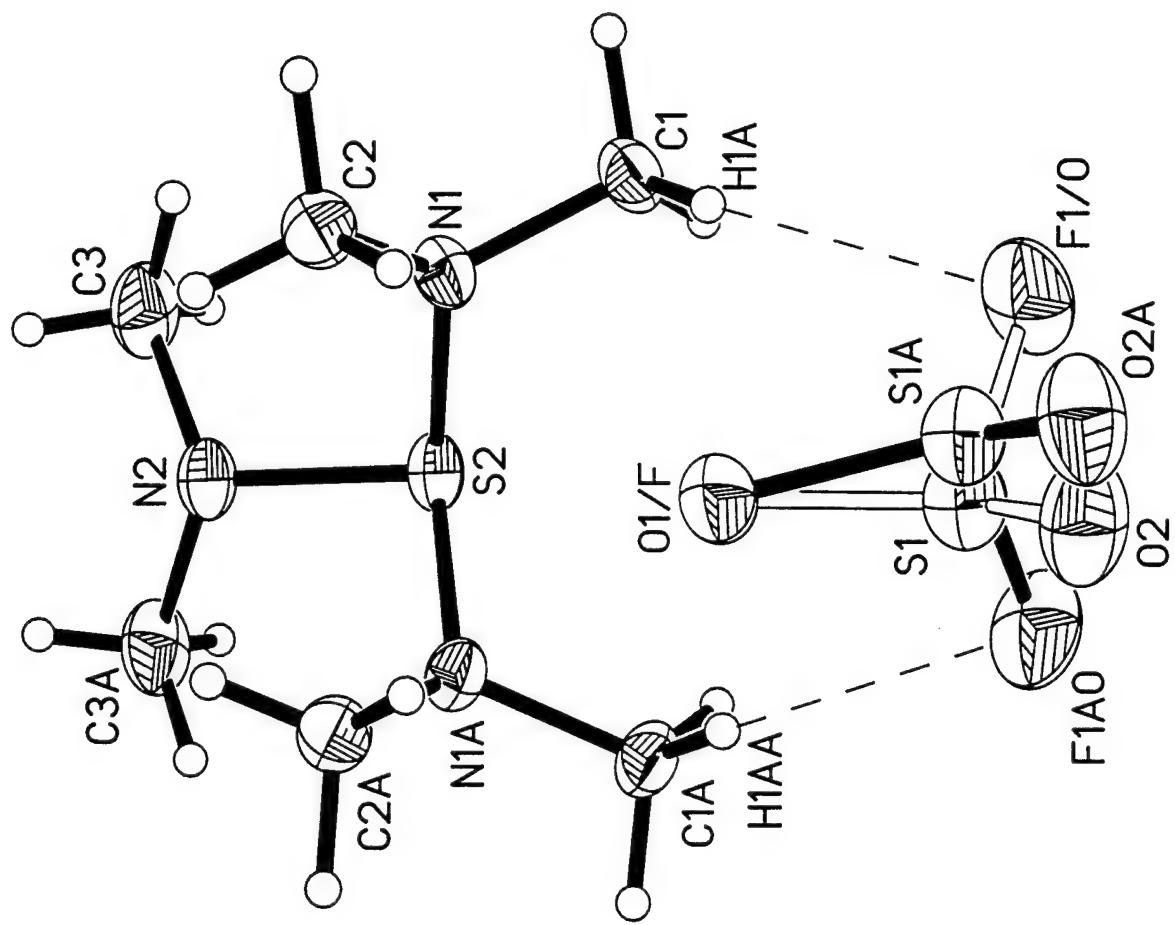
Figure 4: Packing diagram and hydrogen bridging of  $\text{TAOS}^+\text{SO}_2\text{F}^-$ .

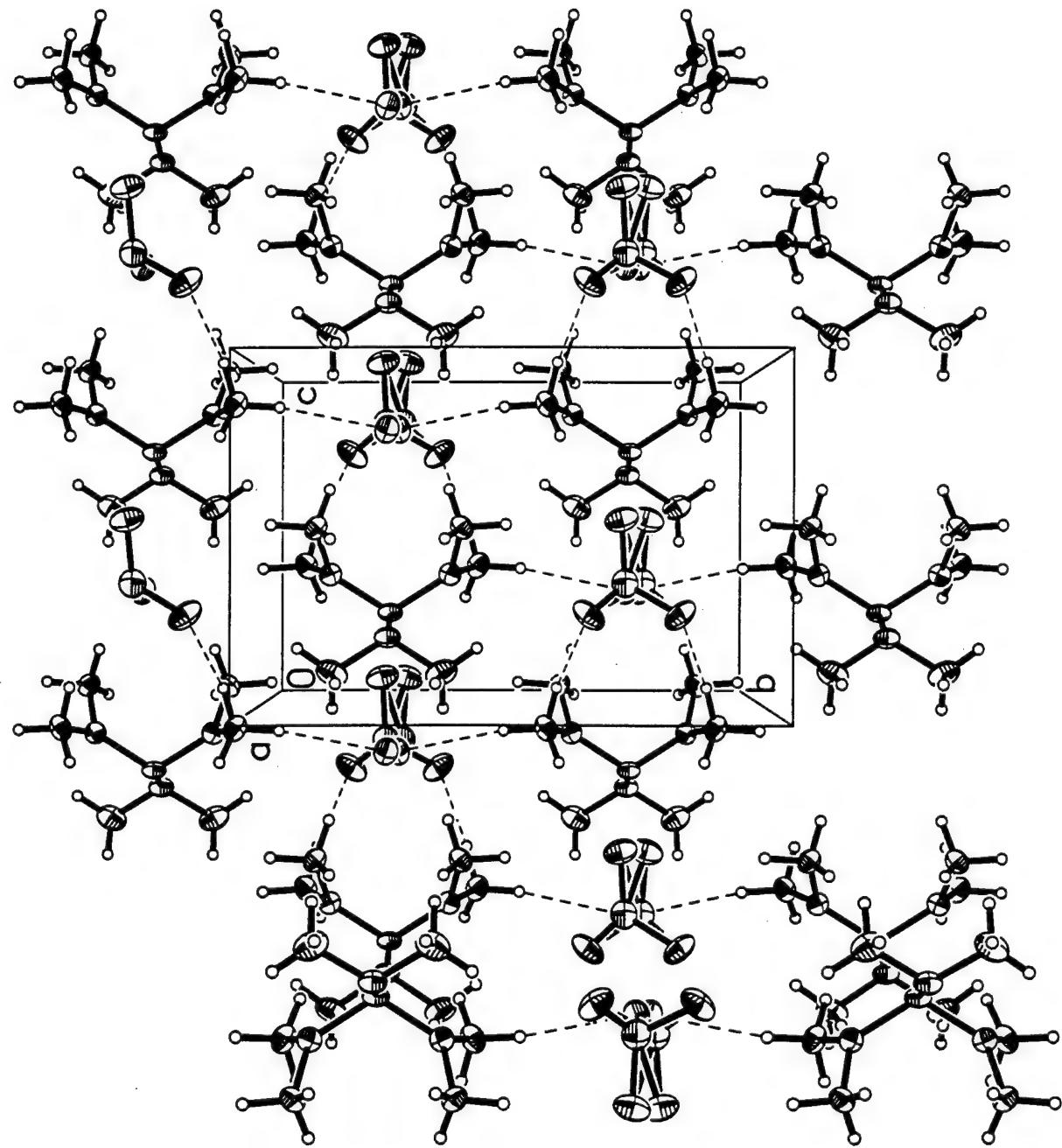
Figure 5: Crystal structure and numbering scheme of  $\text{TMA}^+\text{SO}_2\text{F}^-$  (3). The displacement ellipsoids are drawn at the 50 % probability level, and the hydrogen atoms were located from difference electron density maps.

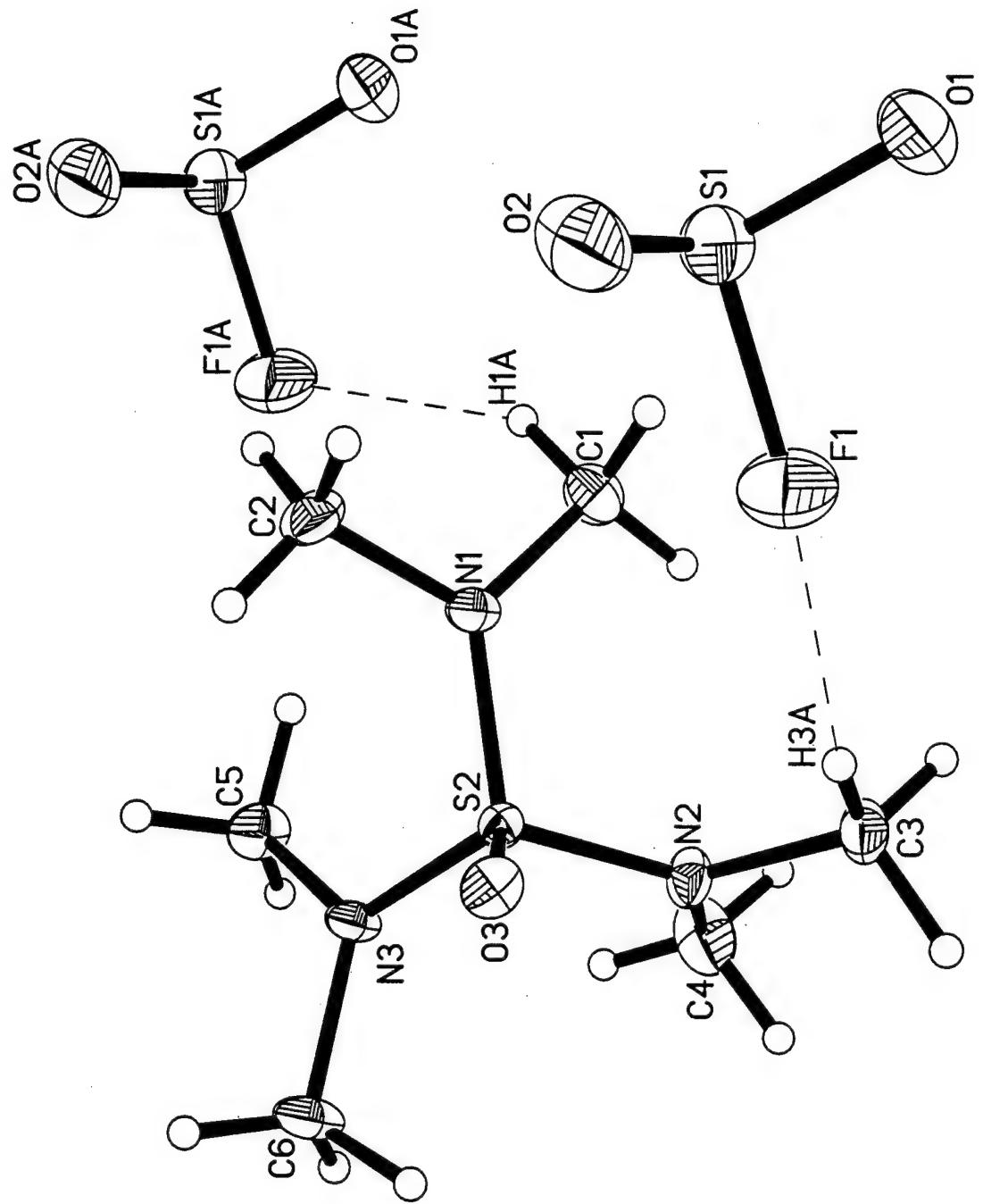
Figure 6: Packing diagram and hydrogen bridging of  $\text{TMA}^+\text{SO}_2\text{F}^-$ .

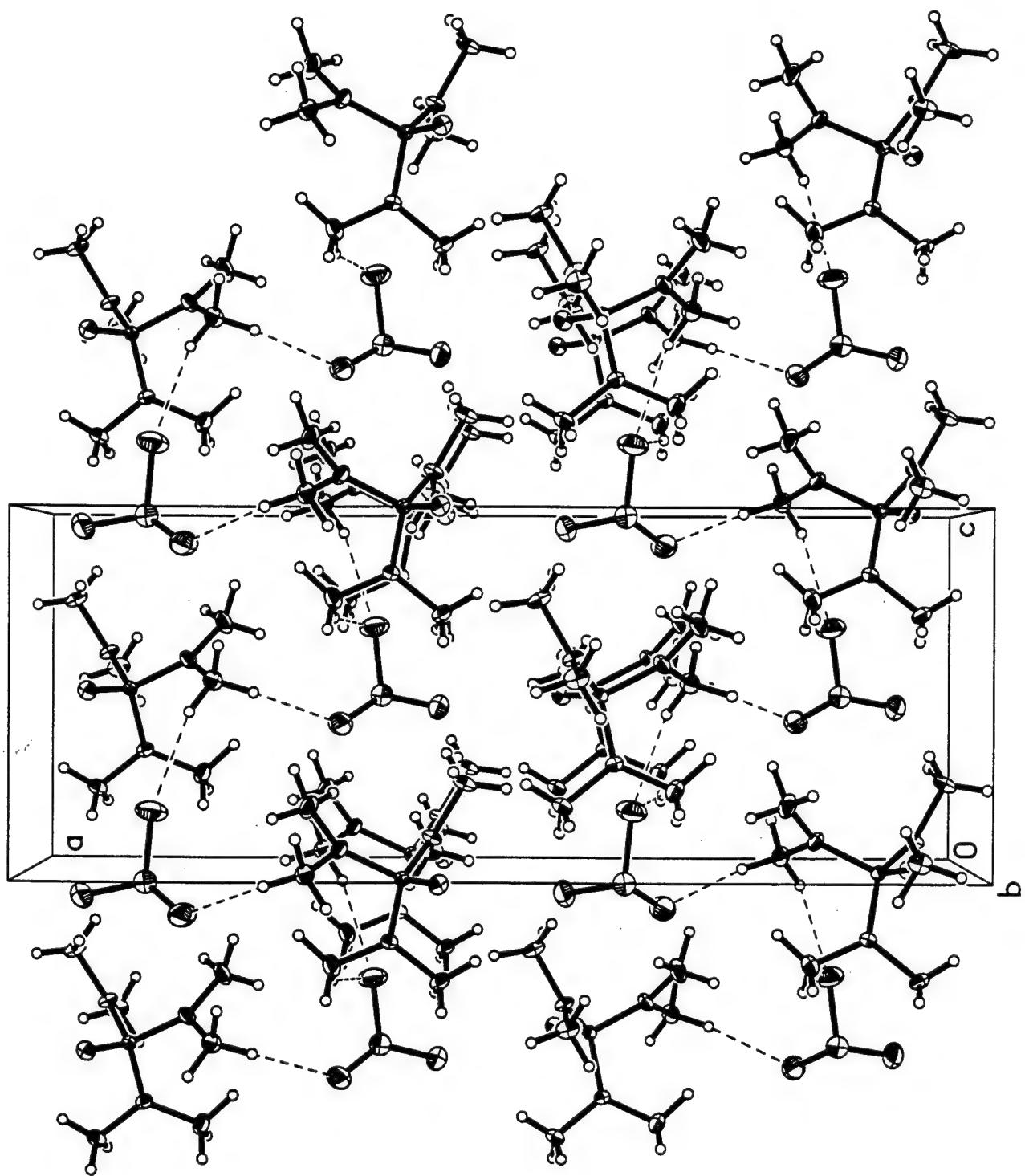
Figure 7: Crystal structure at 113 °K and numbering scheme of  $\text{K}^+\text{SO}_2\text{F}^-$  (4), showing the coordination of 6  $\text{SO}_2\text{F}^-$  anions around the  $\text{K}^+$  cation.

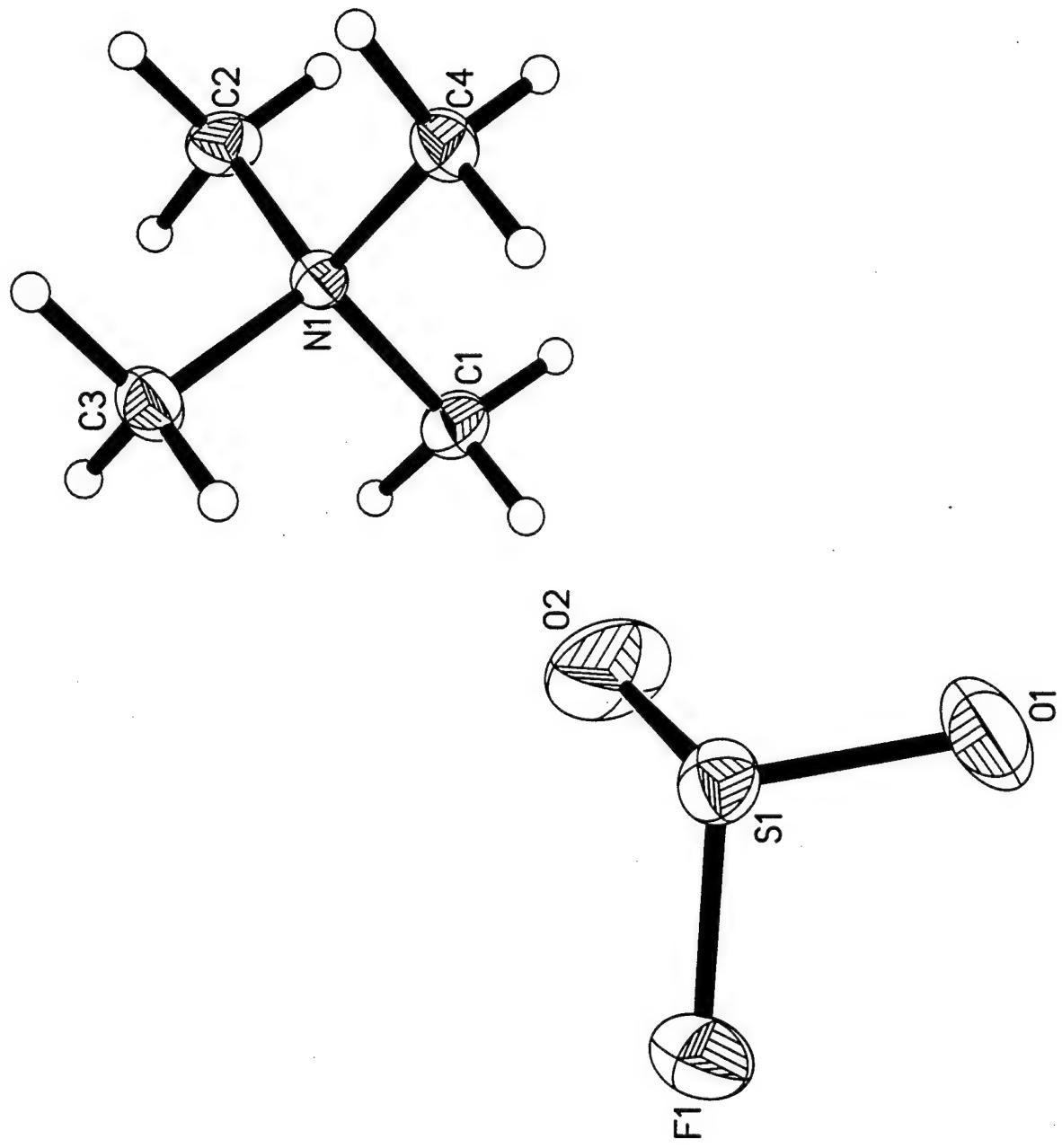
Figure 8: Plot of the observed S-O/F and S-F/O bond lengths of (1)-(4) in pm as a function of their O/F occupancy factors.

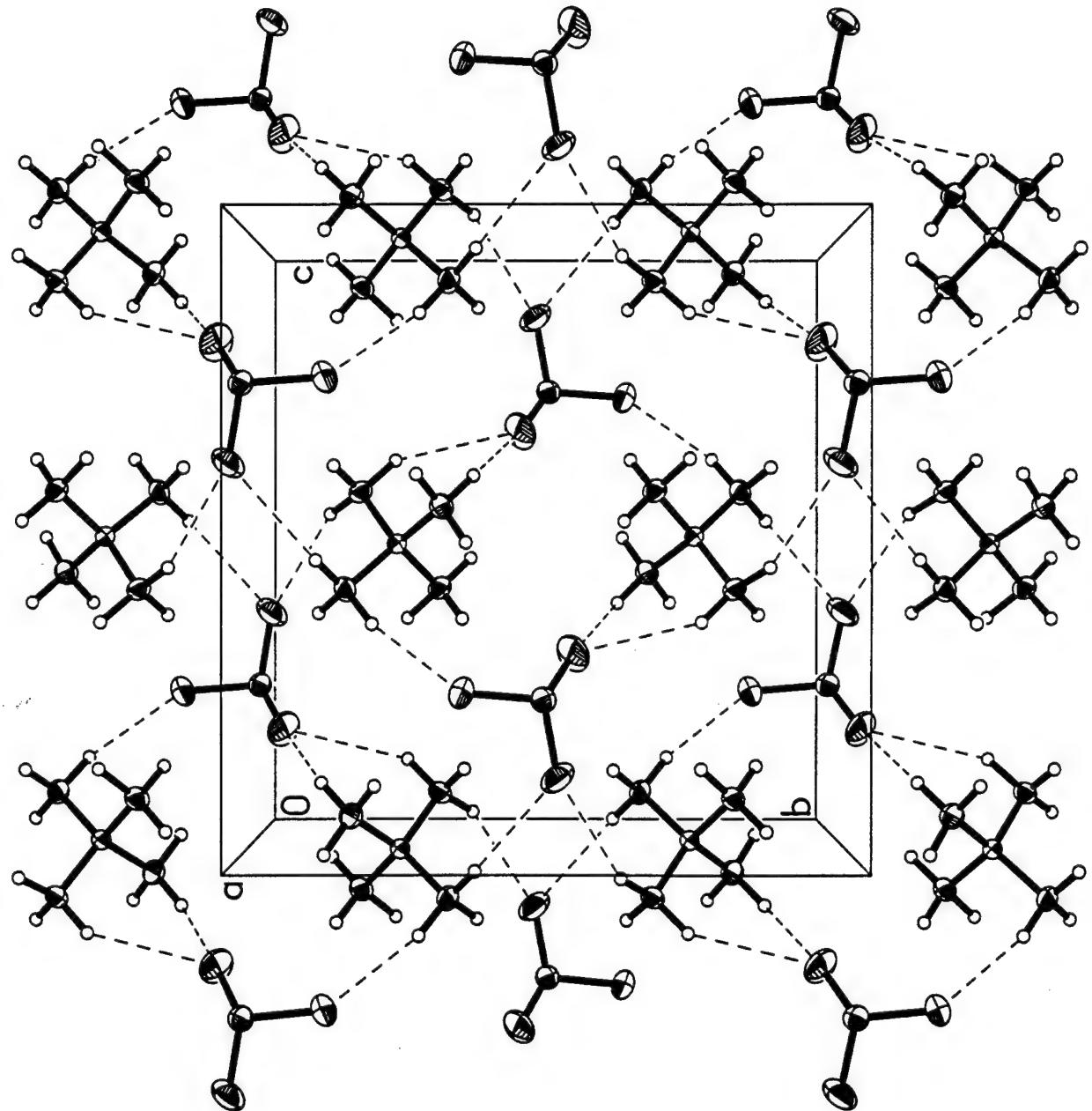


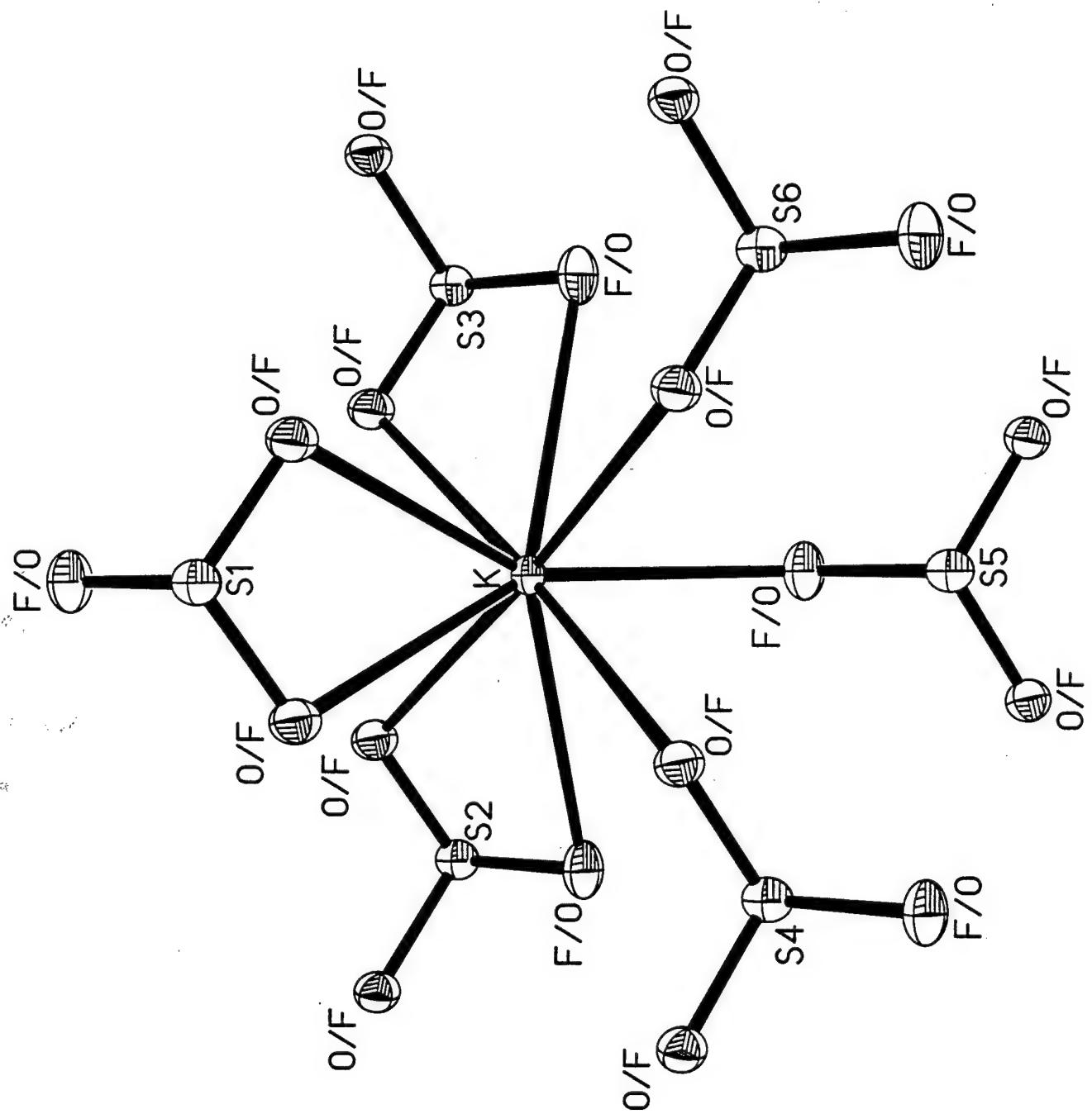












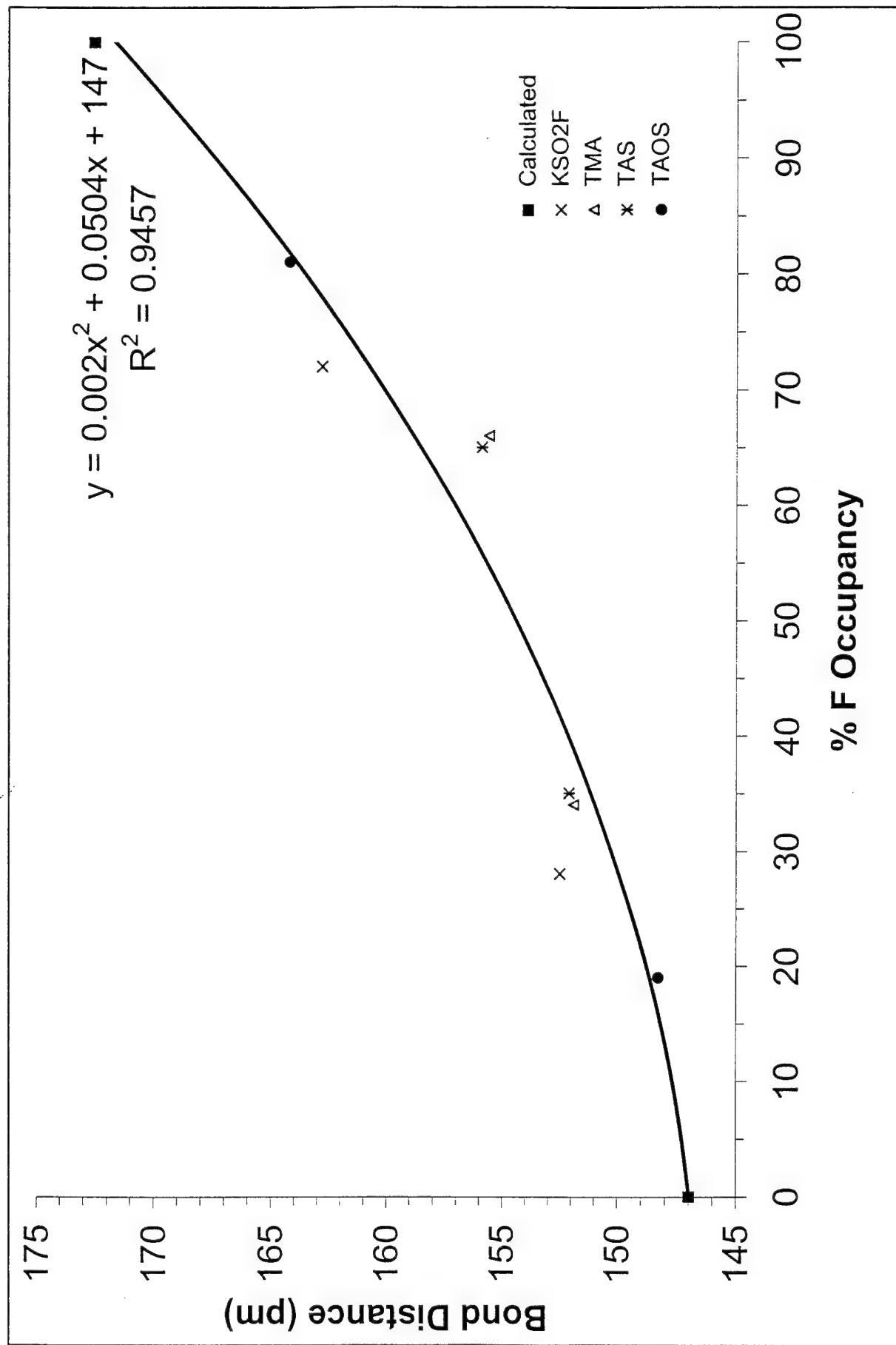


Table 1. Crystal data and structure refinement for (1) - (4)

	1	2	3	4
Empirical formula	$C_6H_{18}FN_3O_2S_2$	$C_6H_{18}FN_3O_3S_2$	$C_4H_{12}FNO_2S$	$FKO_2S$
M	247.35	263.35	157.21	122.16
T (K)	173(2)	113(2)	173(2)	113(2)
Crystal system	orthorhombic	orthorhombic	orthorhombic	monoclinic
Space group	$Pnma$	$Pna2_1$	$Pbca$	$P2_1/m$
a (pm)	1469.0(3)	2185.0(2)	1152.0(1)	461.3(1)
b (pm)	1117.4(2)	673.3(1)	1150.5(1)	564.0(1)
c (pm)	733.4(1)	819.4(1)	1162.7(2)	683.0(1)
$\beta (^{\circ})$				107.06(1)
V(pm)				0.16988(5)
Z	4	4	1.2055(3)	1.5410(3)
$D_c$ (Mgm $^{-3}$ )	1.365	1.451	8	2
Reflections collected	1993	3019	1.355	2.388
Independent	1444	2771	2357	691
R(int)	0.0283	0.0810	1763	427
Parameters	122	0.0468	0.0166	
$\mu(MoK\alpha)$ , mm $^{-1}$	0.438	12	33	
$R_I$	0.0430	0.449	0.375	2.000
$wR_2$	0.0961	0.0434	0.0517	0.0263
Flack's parameter		0.1045	0.1275	0.0748
Difference electron density e. $\text{\AA}^{-3}$	0.412 / -0.257	0.01(9)	0.448 / -0.594	0.600 / -0.407
	0.461 / -0.547			

Table 2. Bond distances and angles of the fluorosulfite anion  $\text{FSO}(1)\text{O}(2)^-$  in different salts compared to those calculated for the free gaseous anion. The data for (5)-(8) are literature values which were not corrected O/F mixed occupancies and therefore, do not represent the true structures.

Compound	SO(1)	SO(2) [pm]	SF	O(1)SO(2)	O(1)SF	Temp. [K] <sup>a</sup>	ref
1 <sup>b</sup>	152.1(3)	146.2(3)	155.9(2)	109.8(2)	108.8(1)	113.0(1)	173
	[O <sub>F</sub> = 0.35]	[O <sub>O</sub> = 1.0]	[O <sub>O</sub> = 0.35]				
2	148.3(3)	146.8(3)	164.2(2)	108.8(2)	102.6(2)	102.4(2)	113
	[O <sub>F</sub> = 0.19]	[O <sub>O</sub> = 1.0]	[O <sub>O</sub> = 0.19]				
3	151.9(2)	147.0(2)	155.6(2)	106.9(2)	102.7(1)	104.2(1)	173
	[O <sub>F</sub> = 0.34]	[O <sub>O</sub> = 1.0]	[O <sub>O</sub> = 0.34]				
4	152.5(2) <sup>f</sup>	152.5(2) <sup>f</sup>	162.8(2)	105.9(1)	101.9(7)	101.9(7) <sup>f</sup>	113
	[O <sub>F</sub> = 0.28]	[O <sub>O</sub> = 0.28]					
Predicted <sup>g</sup> for free gaseous ion	147	147	170-175	113	100.5	100.5	
KSO <sub>2</sub> F (5)	152.6(2)	152.6(2)	159.1(2)	104.9(2)	102.8(1)	102.8(1)	294
RbSO <sub>2</sub> F (6)	153.0(6)	153.6(6)	159.8(8)	105.7(5)	102.6(3)	102.6(3)	294
Ph <sub>3</sub> PCF <sub>2</sub> H <sup>c</sup> (7)	143.6()	141.1(1)	151.6(6)	-	110.5(5)	-	293
F-imid <sup>d</sup> (8)	148.9(11)	146.8(14)	149.6(16)	108.7(7)	104.4(8)	107.4(8)	190
							13,14 <sup>e</sup>

<sup>a</sup>) Temperature of data collection; <sup>b</sup>) The anion is disordered across a mirror plane, the occupancies of mixed O(1) and F positions (O<sub>O</sub> = oxygen occupancy and O<sub>F</sub> = fluorine occupancy) are refined; <sup>c</sup>) Disordered; <sup>d</sup>) F-imid = 2-fluoro-1,3-diisopropyl-4,5-imethylimidazolium cation. <sup>e</sup>) Data uncorrected for mixed O/F occupancies; <sup>f</sup>) Symmetry generated: x, -y+1/2, z; <sup>g</sup>) Predicted values based on the calculations given in Table 3.

**Table 3. Calculated geometries<sup>a</sup> of  $\text{SO}_2\text{F}^-$  compared to the calculated and observed geometries of  $\text{SOF}_2$  and  $\text{SO}_2\text{F}_2$  and the observed geometry of  $\text{ClO}_2\text{F}$**

$\text{SOF}_2$	R1(SO)	R2-R1	R2(SF)	$\angle(\text{FSO})$	$\angle(\text{FSF})$	ref
RHF/6-31 + G*	140.9	16.7	157.6	106.6	92.5	10
RHF/6-311+G(2d)	139.5	15.9	155.4	106.5	92.6	
RHF/aug-cc-pvdz	142.6	16.6	159.2	106.3	92.3	
RHF/aug-cc-pvtz	140.1	14.7	154.8	106.3	93.0	
B3LYP/6-311+G(2d)	142.9	20.4	163.3	106.7	92.8	
B3LYP/aug-cc-pvdz	146.1	20.5	166.6	106.5	92.8	
B3LYP/aug-cc-pvtz	143.4	18.8	162.2	106.4	93.0	
CCSD(T)/6-311+G(2d)	143.1	18.8	161.9	106.7	92.3	
CCSD(T)/aug-cc-pvdz	146.9	19.3	166.2	106.4	92.3	
CCSD(T)/aug-ccpvtz	143.6	16.7	160.3	106.3	92.5	
obsd.	142.0(3)	16.3	158.3(3)	106.2(2)	92.2(2)	28

$\text{SO}_2\text{F}_2$	R1(SO)	R2-R1	R2(SF)	$\angle(\text{OSO})$	$\angle(\text{FSF})$	ref
RHF/6-311 + G(2d)	138.0	12.6	150.6	124.6	95.6	
RHF/aug-cc-pvdz	141.3	13.3	154.6	125.2	94.8	
RHF/aug-cc-pvtz	138.7	11.7	150.4	124.1	95.9	
B3LYP/6-311+G(2d)	141.6	15.7	157.3	125.6	94.9	
B3LYP/aug-cc-pvdz	145.2	16.4	161.6	126.0	94.3	
B3LYP/aug-cc-pvtz	142.2	14.8	157.0	125.2	95.1	
CCSD(T)/6-311+G(2d)	141.6	14.4	156.0	125.5	94.9	
CCSD(T)/aug-cc-pvdz	145.7	15.4	161.1	126.3	94.0	
CCSD(T)/aug-cc-pvtz	142.1	13.3	155.4	125.3	95.1	
obsd.	140.5	12.5	153.0	124.0	96.1	30

$\text{SO}_2\text{F}^-$	R1(SO)	R2-R1	R2(SF)	$\angle(\text{FSO})$	$\angle(\text{OSO})$	ref
RHF/6-31 + G*	145.8	24.0	169.8	100.6	113.2	10
RHF/6-31 + G(3df)	143.7	22.7	166.4	100.5	112.9	14
RHF/6-311 + G(2d)	144.1	19.8	163.9	100.4	113.1	
RHF/aug-cc-pvdz	147.8	23.4	171.2	100.6	112.6	
RHF/aug-cc-pvtz	144.8	21.1	165.9	100.5	112.8	
MP2/6-31+G(3df)	147.1	30.9	178.0	100.3	113.8	14
B3LYP/6-311+G(2d)	147.8	37.2	185.0	100.8	113.4	
B3LYP/aug-cc-pvdz	151.7	32.7	184.4	101.1	112.6	
B3LYP/aug-cc-pvtz	148.3	31.9	180.2	100.7	113.0	
CCSD(T)/6-311+G(2d)	147.9	34.5	182.4	100.5	113.6	
CCSD(T)/aug-cc-pvdz	152.4	30.2	182.6	100.7	112.8	
CCSD(T)/aug-cc-pvtz	148.4	27.3	175.7	100.3	115.9	
Predicted	147	23-29	170-176	100.5	113	

$\text{ClO}_2\text{F}$	R1(ClO)	R2-R1	R2(ClF)	$\angle(\text{FClO})$	$\angle(\text{OClO})$	ref
	142.0	27.4	169.4	101.8	115.2	11

a) Bond lengths in pm and angles in degrees.

Table 4. Calculated scaled (unscaled) and observed vibrational frequencies of  $\text{SO}_2\text{F}_2$ ,  $\text{SOF}_2$ , and  $\text{SO}_2\text{F}$

		$\text{a}_1$			$\text{a}_2$			$\text{b}_1$			$\text{b}_2$			average
$\text{SO}_2\text{F}_2$ ( $\text{C}_{2v}$ )	$\nu_1$	$\nu_2$	$\nu_3$	$\nu_4$	$\nu_5$	$\nu_6$	$\nu_7$	$\nu_8$	$\nu_9$	$\Delta(\nu_{\text{obsd}} - \nu_{\text{calcd}})$	ref			31
Observed	1270	849	553	384	384	1504	544	887	540					
RHF/6-311+G(2d)	1259(1384)	869(955)	557(612)	384(422)	379(417)	1473(1619)	547(595)	897(986)	541(601)	9.4				
RHF/aug-cc-pvdz	1234(1305)	881(932)	561(569)	383(389)	374(380)	1444(1528)	555(551)	913(966)	543(563)	21.5				
RHF/aug-cc-pvtz	1246(1393)	876(979)	550(615)	385(426)	380(421)	1453(1624)	558(601)	904(1010)	542(606)	15.9				
IR and RA Int. <sup>a</sup>	212.5;13.0, .09	128.2;11.5,0	47.9;1.7,73	.12;64.75	0;.99	384.3;1.72	44.7;1.21	276.5;1.81	42.1;1.29					
B3LYP/6-311+G(2d)	1304(1226)	823(774)	557(516)	383(355)	383(355)	1553(1460)	546(506)	859(808)	542(502)	16.3				
B3LYP/aug-cc-pvdz	1281(1162)	841(763)	560(477)	380(324)	378(322)	1518(1377)	551(466)	875(794)	547(469)	8.4				
B3LYP/aug-cc-pvtz	1287(1231)	837(800)	557(520)	384(359)	383(358)	1527(1460)	547(511)	870(832)	541(505)	7.5				
CCSD(T)/6-311+G(2d)	1289(1238)	833(800)	555(532)	387(372)	382(366)	1531(1471)	544(521)	876(841)	538(515)	9.0				
CCSD(T)/aug-cc-pvdz	1262(1159)	848(779)	558(484)	385(334)	377(327)	1498(1376)	548(474)	893(820)	546(475)	4.9				

<sup>a</sup> IR and Raman intensities in  $\text{km mol}^{-1}$  and  $\text{\AA}^4 \text{amu}^{-1}$ , respectively, and polarization ratios.

	$\Delta(\nu_{\text{obsd}} - \nu_{\text{calcd}})$	ref				
$\text{SOF}_2 (C_s)$	$\nu_1$	$\nu_2$	$\nu_3$	$\nu_4$	$\nu_5$	$\nu_6$
Observed	1333	808	530	378	747	393
RHF/6-311+G(2d)	1321(1457)	811(894)	533(601)	380(428)	751(828)	389(439)
RHF/aug-cc-pvdz	1269(1389)	822(900)	536(565)	376(396)	773(846)	391(412)
RHF/aug-cc-pvtz	1288(1466)	819(932)	531(611)	377(434)	764(870)	390(449)
B3LYP/6-311+G(2d)	1420(1313)	794(734)	535(486)	369(335)	716(662)	399(362)
B3LYP/aug-cc-pvdz	1348(1261)	804(752)	540(465)	364(313)	743(695)	401(345)
B3LYP/aug-cc-pvtz	1370(1318)	802(771)	535(501)	370(346)	733(705)	398(373)
CCSD(T)/6-311+G(2d)	1390(1307)	798(750)	532(500)	376(354)	729(685)	392(369)
CCSD(T)/aug-cc-pvdz	1313(1241)	809(765)	536(469)	369(324)	757(716)	399(349)
CCSD(T)/aug-cc-pvtz	1330(1322)	808(803)	531(517)	378(368)	749(745)	392(381)

	$\text{SO}_2\text{F}^- (\text{C}_s)$	$\nu_1$	$\nu_2$	$\nu_3$	$\nu_4$	$\nu_5$	$\nu_6$	$\Delta(\nu_{\text{obsd}} - \nu_{\text{calcd}})$	ref
Observed		1108	590	497	387	1184	363		10
RHF/6-311+G*	1086	605	536	394	1174	362		15.7	
RHF/6-311+G(2d)	1098(1219)	597(663)	511(567)	377(418)	1184(1314)	362(402)		6.8	
RHF/aug-cc-pvdz	1088(1163)	611(653)	529(584)	381(421)	1166(1247)	347(383)		18.8	
RHF/aug-cc-pvtz	1088(1227)	608(686)	514(612)	384(458)	1170(1319)	354(422)		13.5	
B3LYP/6-311+G(2d)	1114(1079)	571(553)	554(472)	364(310)	1218(1179)	350(298)		25.3	
B3LYP/aug-cc-pvdz	1117(1034)	578(535)	558(479)	371(318)	1200(1111)	340(292)		22.8	
B3LYP/aug-cc-pvtz	1113(1086)	577(563)	540(490)	370(336)	1206(1177)	351(318)		18.7	
CCSD(T)/6-311+G(2d)	1105(1076)	576(561)	538(475)	367(324)	1215(1183)	355(313)		19.5	
CCSD(T)/aug-cc-pvdz	1102(1017)	589(543)	550(485)	364(330)	1193(1101)	342(302)		17.2	
CCSD(T)/aug-cc-pvtz	1102(1092)	585(580)	525(506)	378(365)	1201(1190)	352(340)		12.7	

Table 5. Unscaled CCSD(T)/aug-cc-pvtz symmetry force constants and potential energy distribution of  $\text{SO}_2$

Calcd. Freq., cm <sup>-1</sup>		Symmetry force constants <sup>a</sup>				PED <sup>b</sup>
a'	V <sub>1</sub>	1092	F <sub>11</sub>	8.520	F <sub>33</sub>	F <sub>44</sub>
V <sub>2</sub>	580	F <sub>22</sub>	0.486	1.626		
V <sub>3</sub>	506	F <sub>33</sub>	0.220	-0.076	1.913	
V <sub>4</sub>	365	F <sub>44</sub>	-0.007	0.014	0.596	1.924
					F <sub>55</sub>	F <sub>66</sub>
a''	V <sub>5</sub>	1190	F <sub>55</sub>	7.840		
V <sub>6</sub>	340	F <sub>66</sub>	0.100	1.049		
					99.1(5)	0.8(6)
					99.1(6)	0.8(5)
		$f_{\text{iso}} = 8.180$				
		$f_{\text{SF}} = 1.626$				

$$f_{iso} = 8.180$$

$$f_{SF} = 1.626$$

<sup>a</sup> Stretching force constants in mdyn/Å, deformation constants in mdynÅ/rad<sup>2</sup>, and stretch-bend interaction constants in mdyn/rad.  
<sup>b</sup> PED in percent; the symmetry coordinates are defined as: (1)  $V_{\text{sym}}\text{SO}_2$ , (2)  $V\text{SF}$ , (3)  $\delta_{\text{sciss}}\text{SO}_2$ , (4) (3)  $\delta_{\text{sym}}\text{O}_2\text{SF}$ , (5)  $V_{\text{asym}}\text{SO}_2$ , (6)  $\delta_{\text{asym}}\text{O}_2\text{SF}$